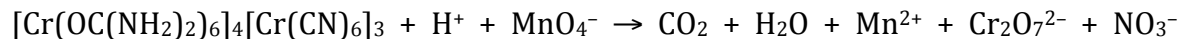
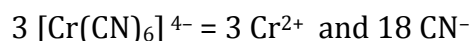
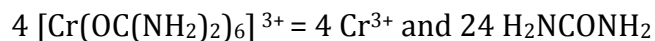


**Balance me! If you can...**

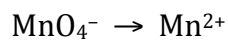
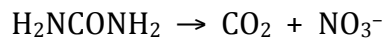
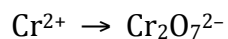
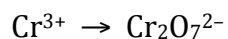


**Hints:**

First deconstruct the chromium double salt into its constituent parts by examining its stoichiometry



Balance the following half-reactions for atoms and charge using water, acid ( $\text{H}^+$ ), and electrons ( $\text{e}^-$ ) taking into account the ratios given above



Add up the half-reaction in multiples that allow the electrons to be canceled. Reassemble  $\text{Cr}^{3+}$ ,  $\text{H}_2\text{NCONH}_2$ ,  $\text{Cr}^{2+}$ , and  $\text{CN}^-$  back into the double salt,  $[\text{Cr}(\text{OC}(\text{NH}_2)_2)_6]_4[\text{Cr}(\text{CN})_6]_3$ .